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NMR investigations of Polytrifluoroethylene (PTrFE) made by RAFT

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Trifluoroethylene (TrFE) is a relatively rare fluorinated monomer mainly used in copolymerisation with vinylidene fluoride (VDF) to prepare ferroelectric materials. While the VDF homopolymerisation has been relatively well studied, that of TrFE is still poorly understood and the reversible deactivation radical polymerisation of this monomer has never been studied in depth. To better understand the RAFT polymerisation of TrFE, the accurate assignments of PTrFE made by RAFT polymerisation is necessary. Thus, this article reports detailed ¹⁹F, ¹H and ¹³C 1D and 2D experiments carried out to determine and assign the different NMR chemical shifts and splitting patterns of the α - and ω -chain ends of PTrFE made by RAFT polymerisation.

A. Introduction

Fluoropolymers constitute an uncommon class of polymers which possess remarkable properties such as a high thermal stability, chemical resistance, good weathering durability, hydrophobic and oleophobic properties as well as electroactivity (ferroelectricity, piezoelectricity...).^{1,2} Within this large family, polyvinylidene fluoride (PVDF) is the second most emblematic fluoropolymers (after polytetrafluoroethylene) and has been studied extensively.³ PVDF is especially exploited for its electroactive properties and his monomer is very often copolymerised with TrFE (trifluoroethylene) affording better processability and crystallizing spontaneously in the electroactive phase without the need to stretch process (in comparison to PVDF). While PVDF and P(VDF-co-TrFE) copolymers are relatively common fluoropolymers, PTrFE is rare and much less studied likely due to his cost and poor availability. Therefore, reports detailing PTrFE microstructure and electroactive properties are very scarce.⁴⁻⁶ The first study of the microstructure of PTrFE was made by Naylor and Lasoski,⁷ who reported assignments of the CF₂ and CFH NMR resonances of the polymer backbone. The work of Wilson *et al.*⁸ provided the first proof of chain defects in PTrFE. Indeed, TrFE undergoes reverse additions during polymerisation: a radical can attack either onto the monomer tail (the CFH group) or onto the head (the CF₂ group). Yagi estimated the amount of backward-added TrFE using Monte Carlo simulations as around 50%.⁹ It is important to note, however, that Yagi considered backward-added monomers, that is to say monomers inserted via head-to-head (HH) and tail-to-head (TH) additions. This author did not consider tail-to-tail (TT) additions leading to regular (head) propagating radicals. In contrast, all the subsequent studies focussed on chain defects defined as HH-TT addition sequences, probably by analogy with PVDF in which an HH addition is systematically followed by a TT addition.¹⁰ Importantly, only Yagi mentioned the possibility of TH additions. Tonelli *et al.*¹¹ made an estimation of the chain defects using ¹³C NMR. They used the Rotational Isomeric State (RIS) model to predict theoretical spectra and better assign the NMR signals. By integrating the ¹³C NMR spectrum of PTrFE they estimated the chain defects to represent 50% of the total number of additions.¹¹ Later, they used the same methodology with ¹⁹F NMR spectra and adjusted their theoretical prediction (leading to a new estimation: 11.6% of chain defects) thanks to the better resolution afforded by ¹⁹F NMR compared to ¹³C NMR.¹² Finally, the same authors reported another allegedly more accurate value for this proportion of chain defects (20%) using ¹H decoupled ¹⁹F NMR spectra.¹³ More recently, Harris

et al. estimated the amount of reverse additions as 13.5% using the signal assignments reported by Cais *et al* and better ¹H decoupling NMR sequence.¹⁴

The homopolymerisation of TrFE and the microstructure and properties of PTrFE have been rarely studied. Although TrFE is mentioned as a comonomer in patents dealing with the RAFT polymerisation of fluoromonomers¹⁵, most PTrFE investigated so far were prepared by conventional radical polymerisation. Reversible deactivation radical polymerisation (RDRP) techniques can help to understand the polymerisation behaviour of fluoromonomers such as VDF or TrFE and to access interesting polymer architectures possessing additional properties. The RAFT polymerisation of VDF has been studied extensively,^{16,17} and a thorough NMR investigation of PVDF made by RAFT gave valuable insights.¹⁸ Nonetheless, the study of the TrFE RAFT polymerisation is more challenging than that of VDF, due to the presence of the CFH stereogenic center and the higher propensity of TrFE for reverse additions. While a fuller investigation of the TrFE RAFT polymerisation is reported and discussed in another publication, the present article describes in details the NMR investigations carried out to identify the PTrFE microstructure and end-groups produced in the course of the RAFT polymerisation of TrFE.

B. Experimental

B.1 Materials

All reagents were used as received unless otherwise stated. Trifluoroethylene (TrFE) was kindly supplied by Arkema (Pierre-Bénite, France). *O*-Ethyl-*S*-(1-methoxycarbonyl)ethylthiocarbonate (CTA-XA) was prepared according to the procedure reported by Liu *et al.*¹⁹ *Tert*-amyl peroxy-2-ethylhexanoate (purity 95%), (Trigonox 121) was purchased from AkzoNobel (Chalons-en-Champagne, France). ReagentPlus grade (purity >99%) dimethyl carbonate (DMC), and tetrahydrofuran (THF) and laboratory reagent grade hexane (purity >95%) were purchased from Sigma-Aldrich and used as received.

B.2 Synthesis

Two techniques were used to synthesize PTrFE.

Reactor procedure. The synthesis of the PTrFE (entry 4, Table 1) used for the NMR analyses was performed by RAFT polymerisation in a 50 mL Hastelloy Parr autoclave system (HC 276), equipped with a mechanical Hastelloy stirring system, a rupture disk (3000 PSI), inlet and outlet valves, and a Parr electronic controller to regulate the

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stirring speed and the heating. Prior to reaction, the autoclave was pressurised with 30 bars of nitrogen to check for leaks. The autoclave was then kept under vacuum (20×10^{-3} bar) for 30 minutes to remove any trace of oxygen. A degassed solution of *tert*-amyl peroxy-2-ethylhexanoate, the initiator (0.281 g, 1.22×10^{-3} mol), and CTA-XA (1.27 g, 6.09×10^{-3} mol) was introduced via a funnel under vacuum. The reactor was then cooled using a liquid nitrogen bath and 10 g of TrFE was transferred by double weighing (i.e. mass difference before and after filling the autoclave with TrFE). After warming to ambient temperature, the autoclave was heated to the target temperature under mechanical stirring. The reaction was stopped after 30 min. The autoclave was cooled to room temperature (ca. 20°C), purged the residual monomer, and the dimethylcarbonate was removed under vacuum. The crude product was dissolved in 10 mL of acetone and left under vigorous stirring for 10 min. This polymer was then precipitated by addition of the acetone solution to 100 mL of chilled hexane. The precipitated polymer (yellow wax) was filtered through a filter funnel and dried under vacuum (15×10^{-3} mbar) for 2 h at 40°C . The polymerisation yield (2.5 % relative to the monomer) was determined gravimetrically.

Carius tube procedure. The TrFE RAFT polymerisation was carried out in thick 8 mL Carius tubes in which a solution of the Trigonox[®] 121 initiator and CTA-XA in DMC (5 mL) was introduced and then degassed by three freeze–pump–thaw cycles. The gaseous monomer was transferred into the Carius tubes at the liquid nitrogen temperature (TrFE, 1.5 g, 1.83×10^{-2} mol, $0.8 \Delta\text{P}$) using a custom-made manifold that enables accurate measurement of quantities of gas (using “pressure drop vs. mass of monomer” calibration curves). The tubes were then sealed under dynamic vacuum at the liquid nitrogen temperature, before being placed horizontally in a shaking thermostatic water bath at 73°C . The tubes were opened at the

desired time and the dimethylcarbonate was removed under vacuum.

The monomer conversions were assumed to be equivalent to the polymer yields, since conversion is very difficult to measure accurately for TrFE or other gaseous monomers.

B.3 NMR spectroscopy

The polymer NMR spectra were collected at 25°C on a Bruker Avance III 400-MHz spectrometer equipped with two independent broadband (^{15}N – ^{31}P and ^{15}N – ^{19}F , 300 W) and a high band (^1H , 100W) rf channels. A 5 mm $^1\text{H}/^{19}\text{F}/^{13}\text{C}$ TXO triple resonance pulsed field gradient probe for which ^{13}C and ^{19}F are on the inner coil and ^1H on the outer coil is used for three channels experiments. This probe has a lower background ^{19}F signals compared to standard dual-channel probes. This triple resonance $^1\text{H}/^{19}\text{F}/^{13}\text{C}$ probe is capable of producing short 90° pulses of $6.5 \mu\text{s}$ width for ^{19}F , $9.5 \mu\text{s}$ for ^{13}C and $9.2 \mu\text{s}$ for ^1H channels. In all experiments, ^1H decoupling is realized with waltz16. ^{19}F decoupling was performed with nested loops using 0.5 ms and 1 ms chirped adiabatic pulses with 80 kHz bandwidth in order to desynchronize and minimize decoupling artifacts.

^1H 1D NMR. A one pulse 90° ($9.25 \mu\text{s}$) pulse sequence was used with 6 s acquisition time, 3 kHz spectral window, 1 transient and 1s recycle delay.

^{19}F 1D NMR. A one pulse 90° ($6.5 \mu\text{s}$) pulse sequence was used with 0.08 s acquisition time, 75 kHz spectral window 1 transient and 1s recycle delay.

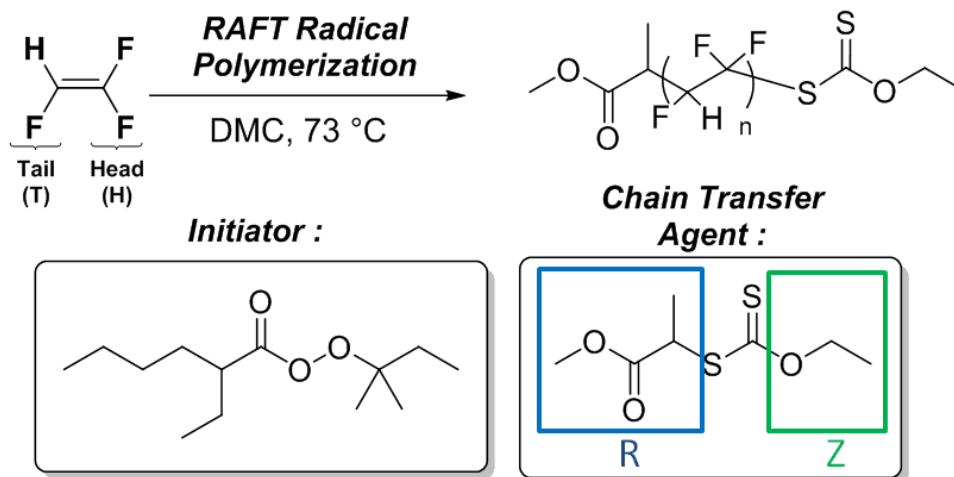


Figure 1. Scheme of the RAFT polymerization of TrFE, using *tert*-amyl peroxy-2-ethylhexanoate as initiator, *O*-ethyl-S-(1-methoxycarbonyl)ethyl)dithiocarbonate as CTA and dimethylcarbonate as solvent.

Table 1. Experimental Conditions and Results for the RAFT Polymerization of TrFE

Entry	[TrFE] ₀ : [CTA] ₀ : [I] ₀	Polymerization time (h)	Conversion (%)	M_n^1 (g/mol)	\bar{D}	CFH-XA ² (%)	CF ₂ -XA ² (%)	Irreversible transfer ² (%)
1	100:1:0.2	2	4.3	2,100	1.06	90.9	9.1	0
2	100:1:0.2	15	67	9,900	1.58	32.4	0	67.6
3	50:0:0.2	15	79	5,700	4.30	0	0	100
4	20:1:0.1	0.5	2.5	n.d.	n.d.	n.d.	n.d.	n.d.

Note: n.a. = not applicable, polymer made by conventional radical polymerization in solution.

n.d. = not determined. Due to insufficient quantity of polymer recovered after polymerization, no SEC measurement was done on this sample. ¹Calculated through GPC using PMMA standards. ²Calculated using NMR and equations S7 to S10.

The theoretical DP of the PTrFE of entry 4 is 0.5, however, estimation of the actual DP by ^1H and ^{19}F NMR indicate a DP of 7. These results suggest that only 7 % of the CTA actually took part in the polymerization at that stage.

¹³C 1D NMR with ¹H, ¹⁹F and ¹H+¹⁹F Decoupling. A one pulse 90° pulse sequence was used with 1.1 s acquisition time, 30 kHz spectral window 4100 transient and 1 s recycle delay.

¹⁹F 2D NMR COSY with ¹H Decoupling. The cosygp pulse sequence from Bruker catalog was modified in order to include ¹H decoupling over the whole pulse sequence. The acquisition parameters were 1 s acquisition time, 75 kHz spectral windows in F2 and in F1, 4 transients and recycle delay of 1 s. Processing involved a magnitude calculation phase correction in the F1 dimension.

¹H{¹³C} 2D NMR HSQC with ¹H Decoupling. The hsqcetgpsi2 HSQC pulse sequence from Bruker catalog was modified in order to apply ¹⁹F decoupling over the whole pulse sequence. Acquisition parameters were 0.3 s acquisition time, 7.5 kHz spectral window in F2, 25 ms acquisition time, 30.2 kHz spectral window in F1, ¹J_{CH} = 152 Hz, garp decoupling for ¹³C, 8 transients and recycle delay of 1 s. Processing involved an exponential window multiplication in both dimensions.

¹⁹F{¹³C} 2D NMR HSQC with ¹H Decoupling. The pulse sequence described by Li et al.²⁰ (2D NMR studies of a model for Krytox® perfluoropolyethers) was written from scratch for a Bruker system, the only modifications being ¹H decoupling over the whole pulse sequence, ¹³C decoupling performed with nested loops using 0.5 ms and 1 ms chirped adiabatic pulses with 30 kHz bandwidth in order to desynchronize and minimize decoupling artefacts and echo-antiecho quadrature detection in F1. Acquisition parameters were 83 ms acquisition time and 75 kHz spectral window in F2, 99 ms acquisition time and 10 kHz spectral window in F1, ¹J_{CF} = 260 Hz, 16 transients and recycle delay of 1 s.

Processing involved linear prediction of an exponential window multiplication in both dimensions and a magnitude calculation phase correction in the F1 dimension.

²J_{CF} = 30 Hz was used for the ²J ¹⁹F {¹³C} 2D NMR HSQC with ¹H Decoupling.

C. Results and discussion

Investigating the ¹⁹F NMR of PTrFE made by RAFT is challenging for two reasons: tacticity and chain defects (reverse additions). PTrFE possesses CFH (T) and CF₂ (H) groups producing resonances in two different zones. The CF₂ resonances are found at chemical shifts ranging from -110 to -140 ppm while those of the CFH signals are in the -200 to -220 ppm region. Moreover, the stereogenic CFH groups often split the signals of the neighbouring atoms. For example, the configuration of the two CFH groups in a -CFH-CF₂-CFH- sequence will significantly influence the CF₂ signals. An isotactic triad will produce a *meso* configuration for the central CF₂ with two non-equivalent fluorine atoms (F_a and F_b) and the ¹⁹F{¹H} NMR pattern of the CF₂ group will be an AB system with a strong ²J_{FF} geminal coupling (Figure S1). In contrast, a syndiotactic structure will generate a racemic configuration for the CF₂ group where the fluorine atoms (F_c) are isochronous and their signal will appear as a broad singlet since the ³J_{FF} coupling (from 5 to 30 Hz) on the backbone polymer signal is too small relative to the resonance linewidth (Figure S1). This kind of splitting pattern has been studied and described in previous reports and explains the fact that the resonances of both CF₂ and CFH groups lead to complex signals.^{14,21} This article will rather focus on the RAFT chain ends brought by the RAFT process.

Another complexity of the ¹⁹F NMR spectrum arises from the reverse additions that occur during polymerisation (see Introduction).²¹ The different macroradicals resulting from the regular and reverse monomer additions are described in Figure S2. The head-to-head (HH) additions are minor (13.5% according to

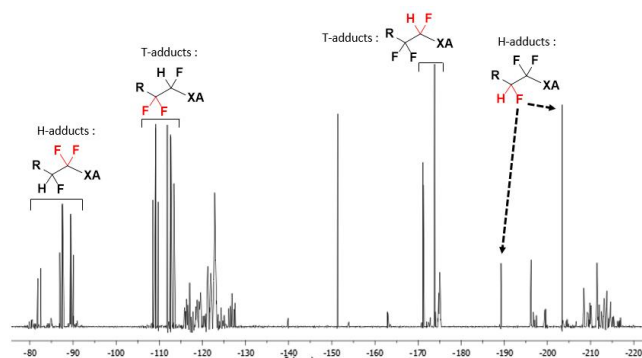


Figure 2. ¹⁹F{¹H} spectrum of PTrFE (yield = 4.3%) made by RAFT (Entry 1, Table 1) recorded in acetone-d₆ (where XA stands for SC(S)OEt and R for CH₃O(C=O)CH(CH₃)).

Tonelli et al¹³) but play an important role on the PTrFE properties (e.g. crystallinity and electroactivity). This article deals with the assignment of the different chain-end groups of PTrFE synthesized using RAFT polymerisation. A study of the TrFE RAFT polymerisation mechanism will be reported elsewhere.²²

The complete assignments of the chain-end signals for the PTrFE product made by RAFT was made possible by a set of 1D and 2D ¹H, ¹⁹F and ¹³C NMR experiments.

C.1 Monoadducts determination

The ¹⁹F{¹H} NMR spectrum (Figure 2) of a PTrFE made by RAFT polymerisation stopped at low conversion (Entry 1, Table 1, 4.3%) shows numerous resonances, and in particular several groups of sharp peaks. These signals reveal the presence of two types of monoadducts (i.e. product resulting from the reaction of only one TrFE unit with the R' radical of the CTA followed by transfer to another CTA). Several isomers can be noted. The first monoadduct results from the R' addition to the TrFE tail (regular addition leading to R-CHF-CF₂-XA, called H-adduct), whereas the second one results from the R' addition to the TrFE head (leading to R-CF₂-CFH-XA and called T-adduct). The possible structures of these monoadducts are shown in Figure S3. The two stereocenters (R is chiral) on the molecules produce four different stereoisomers per type of monoadduct (two enantiomeric pairs of diastereoisomers) with relatively complex NMR splitting patterns.

The monoadduct resonances are easily recognizable as they appear as sharp peaks, contrarily to the broad polymer signals. In addition, the presence of the xanthate end-group strongly shifts the resonances downfield as reported earlier.^{16,23} Consequently, the CF₂ signals of the H-adducts are found at around -80 to -90 ppm while those of the T-adducts are located between -105 and -115 ppm (Figure 2). Similarly, the CFH signals of the H- and T-adducts are located between -185 and -205 ppm and between -170 and -175 ppm, respectively. All of these regions contain two sets of resonances, each one corresponding to one diastereoisomer. The CFH resonance integral allows the determination of each diastereoisomer pair relative proportion (Equations S1 to S4). For both adduct families, the most abundant diastereoisomer should correspond to the RR/SS enantiomeric pair as they present the methyl and fluorine atom in anti-positions (whereas they are syn in the RS/SR couples). This consideration was used to assign the NMR signals to the different diastereoisomers (see Table 2, and Figure S4 to Figure S11).

The 3-bond correlations between the CFH and CF₂ fluorine atoms visible on the ¹⁹F COSY spectrum (Figure S12 and Figure S13) allowed the assignment of each signal to a unique enantiomer couple (see Table 2). The presence of the CFH stereocentre makes both fluorine atoms in the CF₂ group (F₁ and F₂, Table 2) non-equivalent, yielding an ABX pattern (i.e. doublets of doublets for each F atom, in Figure

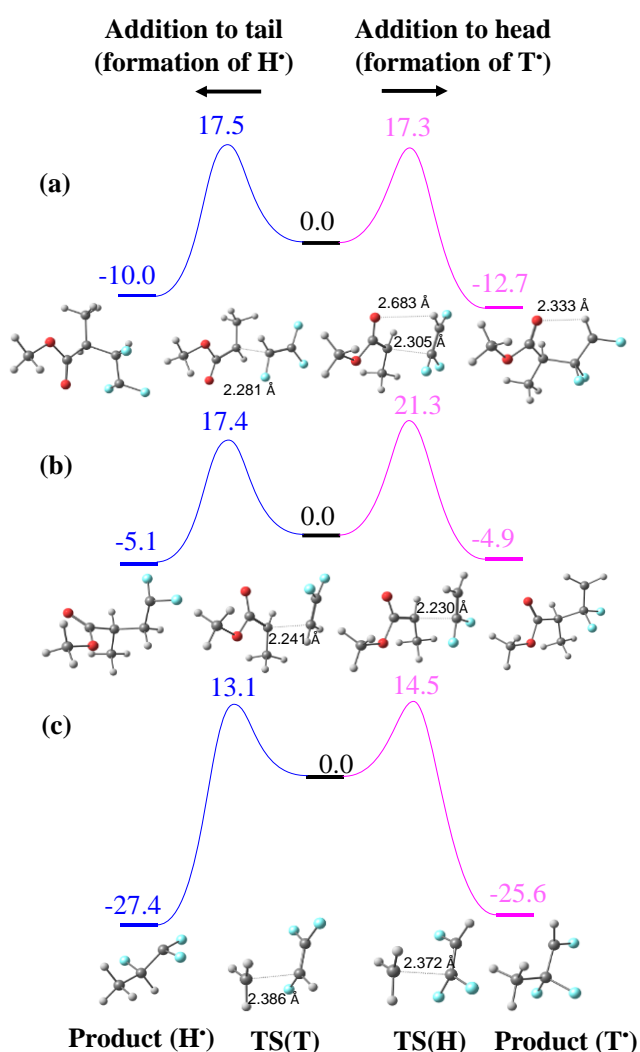


Figure 3: Energy profiles (values are standard free energies at 298K in kcal mol⁻¹) and views of the optimized structures for the starting compounds, transition states and products of radical additions to monomers. (a) Addition of [•]CH(CH₃)COOCH₃ to CHF=CF₂. (b) Addition of [•]CH(CH₃)COOCH₃ to CH₂=CF₂. (c) Addition of CH₃[•] to CHF=CF₂.

S4 to Figure S7 for the H-adducts and Figure S8 to Figure S11 for the T-adducts). The geminal F atoms are strongly coupled, with ²J_{F-F} around 236-237 Hz and 260-263 Hz for the H-adduct (Figures S4 and S6) and T-adduct (Figures S8 and S11), respectively. The CFH

resonance is also complex due to the ³J_{F-F} couplings with F₁ and F₂ (see Figures S5 and S7 for the H-adducts; Figures S9 and S10 for the T-adducts). The difference of the ³J_{F-F} values in the 2 diastereoisomer couples (Table 2) is worth noting. This is ascribed to the different shielding effect of the CFH and CH₃ groups on F₁ and F₂. In the RS/SR diastereoisomers (for both H- and T-adducts), both CFH and CH₃ are *syn*, resulting in a stronger influence on F₁ or F₂ compared to the RR/SS diastereoisomers. Finally, a 2D ¹H - ¹⁹F heterocoupling experiment (Figures S14-S16) allowed us to assign the ¹H NMR resonances of the CFH, CH and CH₃ groups. To confirm the above assignments, the chemical shifts, ²J_{F-F} and ³J_{F-F} constants determined from the ¹⁹F and ¹H spectra (Table 2) were used in the simulation software gNMR and the ¹⁹F{¹H} spectra were reconstructed. The simulated and experimental spectra showed good agreement (Figure S17 and S18), confirming the reliability of the ABX system solving method.²⁴ Surprisingly, the T-adducts seemed more abundant than the H-adducts (Table 2, Equation S5 and S6). This may be partly explained by the slower T-adduct reactivation. Moreover, the addition of the CTAXA R' radical to the TrFE tail or head is not equiprobable. In the present case, the acrylic radical derived from CTAXA may also react preferentially to the TrFE head (leading to T-adducts). In order to cast light on this question, the barriers for the addition of the acrylate radical from CTAXA to both monomer ends were calculated using DFT. The chosen computational level is identical to that used in the previous investigations of both propagation (HT, HH, TH and TT additions) and chain transfer (degenerate H-H and T-T and non-degenerate H-T) in the RAFT polymerisations of VDF¹⁸ (see computational details in the supporting information).

The results (Figure 3a) confirm that the addition to the TrFE head end is more favored: at 25 °C, the free energy of activation is 0.2 kcal mol⁻¹ lower than for the corresponding addition to the tail end. At 70 °C (temperature used for the polymerisation), both barriers are slightly higher (18.9 vs. 19.1 kcal mol⁻¹, respectively) because of a negative activation entropy but a 0.2 kcal mol⁻¹ bias in favor of the head end attack remains. On the basis of this Δ(ΔG[‡]), the expected H/T ratio at 70 °C is 57:43. In order to reproduce the observed H/T ratio (79:21), the Δ(ΔG[‡]) should be ~0.4 kcal mol⁻¹ and we consider this agreement as satisfactory. The most interesting point is the reason leading to this preference, because attack of the two TrFE carbon atoms by other radicals (e.g. the head and tail radicals of the growing PTrFE chain) occurs preferentially at the tail end)

The reason of this discrepancy is clear when viewing the optimized geometries of the transition state and product for the two pathways. Whereas the tail-end attack does not display any interaction other than the incipient C...C bond, the head-end attack also shows a short

Table 2 : Chemical shifts and coupling constants of the H-adduct and T-adduct of TrFE and *O*-ethyl-*S*-(1-methoxycarbonyl)ethylthiocarbonate (CTA-XA).

		Molar fraction (%)	δ CFH (ppm)		δ CF ₂ (ppm)		² J _{F-F} (Hz)	³ J _{F-F} (Hz)	
			H	F	F ₁	F ₂		With F ₁	With F ₂
H-adducts (29 mol%)	RR/SS	69	CFH : 5.53 CH : 3.14 CH ₃ : 1.33	-203.3	-87.2	-89.7	237	16	16.4
	RS/SR	31	CFH : 5.33 CH : 3.14 CH ₃ : n.d.	-189.1	-82.0	-89.6	236	15.5	17.2
T-adducts (71 mol%)	RR/SS	56	CFH : 7.13 CH : 3.48 CH ₃ : 1.41	-173.8	-109.3	-112.9	263	17.7	16.9
	RS/SR	44	CFH : 7.22 CH : n.d. CH ₃ : 1.41	-171.1	-108.8	-112.1	260	17.2	15.5

Note: n.d. = non determined. The corresponding correlation spot on the 2D ¹H COSY is hidden by signal of the α-chain end signal. Nonetheless, the chemical shifts can be estimated to be around 1.3-1.4 ppm for the CH₃ and 3.4-3.5 ppm for the CH group.

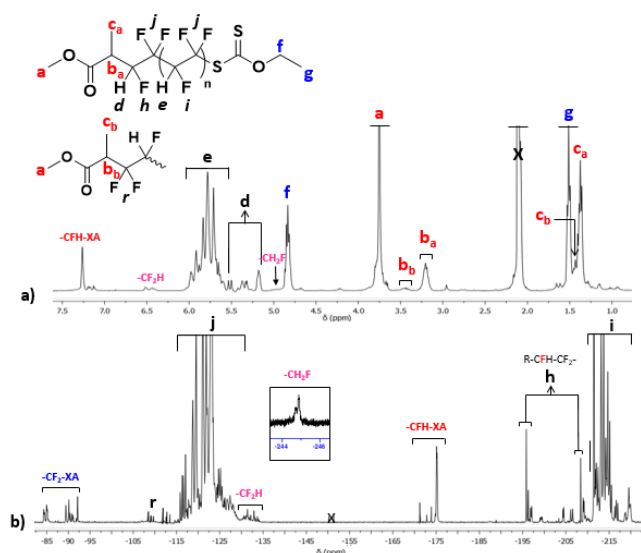


Figure 4. a) $^1\text{H}\{^{19}\text{F}\}\{^{13}\text{C}\}$ and b) $^{19}\text{F}\{^1\text{H}\}\{^{13}\text{C}\}$ spectra in acetone- d_6 of PTTrFE made by RAFT (Entry 4, Table 1).

contact (H-bond) between the H atom of the monomer CHF group and the carbonyl oxygen atom of the attacking radical. This H-bond is already established at the TS level (2.683 Å) and is further strengthened in the product (2.333 Å). This interaction obviously stabilizes the molecule and thus introduces a bias in favor of the head-end attack. For the tail-end addition pathway, on the other hand, the H atom is unsuitably placed to establish a strong H-bond with the carbonyl oxygen atom because the resulting ring would be too strained. A relatively strong H-bond in this system is not surprising when considering the electron withdrawing power of the three F atoms, significantly reducing the electron density on the monomer H atom. In order to further support this rationalization, additional calculations were carried out on another two initiating systems. The first one (Figure 3b) is the addition of the same acrylate radical to VDF, for which $\text{R}_0\text{-CF}_2\text{CH}_2^*$ (product of the addition to the head-end) lacks a $\beta\text{-F}$ substituent and hence the C-H bonds should display lower tendency to act as proton donors in H-bonding. Previous calculations on the VDF RAFT polymerisation were limited to the exchange of head and tail model radical on the CTA, as well as on the regular and reverse monomer additions to the same model radicals¹⁸, but did not address the addition of the CTA primary radical to the VDF monomer. The second initiating system (Figure 3c)

consists of the addition to TrFE of the simpler methyl radical, which lacks the carbonyl function. The results for both systems show that the preferred addition is indeed, as expected, to the monomer tail-end. In particular, a comparison between Figures 3a and 3b shows that the barrier to the tail-end additions (CHF for TrFE and CH_2 for VDF) are essentially identical, whereas the addition to the head CF_2 group requires a much greater barrier for VDF (21.3 kcal mol⁻¹, i.e. 4.0 kcal mol⁻¹ higher than for TrFE). This preferential CTA primary radical addition to the VDF monomer tail was in fact experimentally confirmed by our group in a recent publication.¹⁷ The barriers for the addition of CH_3^* to TrFE are lower than those for the addition of $\text{CH}_3\text{C}^*\text{COOCH}_3$, in agreement with the greater reactivity of the CH_3^* radical, with a difference of 1.4 kcal mol⁻¹ in favour of the tail-end addition and lead to greater thermodynamic gains.

C.2 Assignment of the α -chain-end

There are five possible structures for the α -chain-end (Figure S19). Chains can be initiated by radicals derived from the *tert*-amyl peroxy-2-ethylhexanoate thermal decomposition, by the DMC solvent after H-atom abstraction by the very reactive PTTrFE^{*} radical, or by the CTA R-group. The asymmetric initiator can produce two different radicals and the carboxylate radical may undergo decarboxylation depending on the temperature. The initiator-generated chain ends could be attributed to the low intensity resonances observed between 0.85 and 1.7 ppm and partly overlapping with the signals of the CTA R-group. This overlap prevented their quantification.

Moreover, as mentioned above, the initiating radical can add to either the tail (CHF, noted R-PTTrFE_T) or the head of TrFE (CF_2 noted R-PTTrFE_H). In addition to this regioisomerism, as in the case of monoadducts, the presence of asymmetric carbons results in the generation of pairs of enantiomers and diastereoisomers. The number of stereoisomers is higher than in the case of the monoadducts and it is function of the number of stereocenters (n): (2^n) (Figure S20).

Even with the use of ^{19}F decoupling, the ^1H resonances were too broad to clearly distinguish each diastereoisomer splitting pattern (Figure 4a). Nevertheless, the CTAXA R-group methyl resonance (c_a , Figure 4a) was assigned to the signal at 1.37 ppm in the $^1\text{H}\{^{19}\text{F}\}$ spectrum. The COSY $^1\text{H}\{^{19}\text{F}\}$ spectrum (Figure S21) shows clear correlations of this signal with that of the methanetriyl group (b) at 3.2 ppm, and with that of the CFH moiety (c_{1-4}) 4 bonds away. This 4-bond correlation system is observed as 4 spots (c_{1-4} , Figure S21)

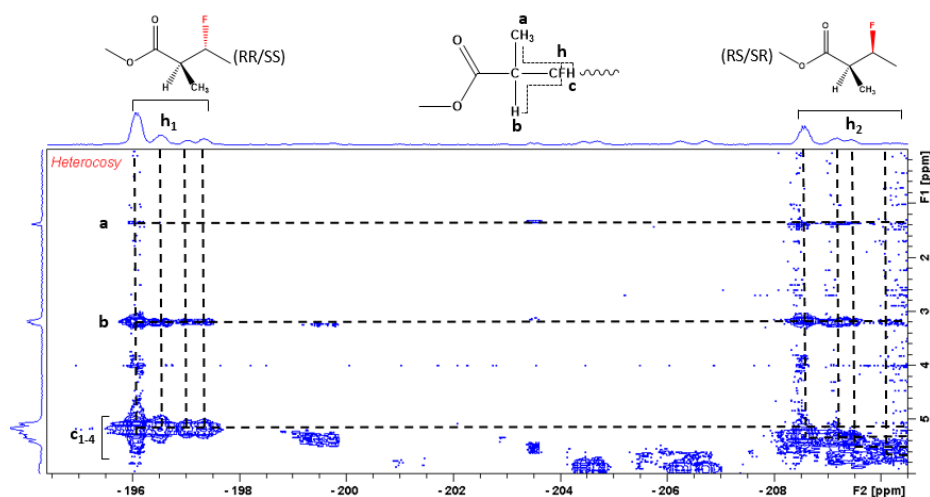


Figure 5. Zoom of the [0 — 6ppm; -194.5 — -210.5ppm] region (RT-PTTrFE correlation zone) of the ^{19}F - ^1H Heterocosity spectrum of PTTrFE made by RAFT recorded in acetone- d_6 (Entry 4, Table 1).

Table 3. Assignments of the ^1H and ^{19}F signals of the R-PTrFE $_T$ α -chain end.

	^1H (ppm)			^{19}F (ppm)
	CH ₃ (a)	CH (b)	CFH (c)	CFH (h)
RR;SS	1.37	3.20	5.17	-196.1
	1.37	3.20	5.18	-196.5
	1.37	3.20	5.18	-197.0
	1.37	3.20	5.18	-197.3
RS; SR	1.37	3.20	5.31	-208.6
	1.37	3.20	5.35	-209.1
	1.37	3.20	5.37	-209.5
	1.37	3.20	5.53	-210.2

between 5.16 and 5.51 ppm corresponding to the 4 couples of enantiomers generated by the first 3 stereocentres (Figure S20).

The ^1H - ^{19}F heterocosity spectrum (Figure 5) shows two groups of signals for the CFH fluorine atom (h_1 and h_2 around -197 and -209 ppm respectively) corresponding to 2 pairs of enantiomers, and correlating through 2, 3 and 4 bonds with the c_{1,c_4} , b and a protons (around 5.3, 3.2 and 3.37 ppm, respectively).

The integration of these fluorine signals revealed h_1 as the signal of the most abundant stereoisomer (60.2%). It was thus assigned to the RR/SS enantiomer couple following the reasoning used previously for

the assignment of the monoadduct signals. It is important to point out that the relative abundance of the RR/SS and RS/SR structures of this R-PTrFE $_T$ end-group is in good agreement with the corresponding relative abundance determined for the T-adduct.

Four resonances were found for each of these groups of fluorine signals (Figure 5) indicating the existence of 8 different enantiomer couples. This observation was confirmed by the ^{19}F - ^{13}C gHSQC spectrum (Figure S22 and Figure S23) which shows 4 different carbon resonances for each fluorine signals. These relatively complex NMR signals can only be explained by considering 4 stereocentres, which underlines the strong influence of stereochemical configurations on the PTrFE ^{19}F NMR signals. Table 3 summarises the chemical shifts and assignments of the ^1H and ^{19}F NMR signals of these PTrFE α -chain end.

C.3 Assignment of the ω -chain-end

Two types of characteristic ω -chain-ends can be produced: dead chains arising from irreversible transfer and dormant chains issued from RAFT polymerisation. Dead chains generated by irreversible transfer (H-abstraction) reactions (to DMC, monomer, or polymer) lead to -CF₂H or -CFH₂ end groups. The ^1H NMR signal (broad triplet at 6.5 ppm, $^2J_{\text{H,F}} = 48$ Hz, Figure 4) of -CF₂H end-group was already reported by Soulestin et al.²⁵ as these chain ends, which are not specific to RAFT polymerisation, are also formed during conventional radical polymerisation. 2D experiments (Heterocosity and gHSQC)

Table 4. Assignments of the ^1H and ^{19}F NMR signals of the R-PTrFE $_T$ α -chain end

CFH-XA						
	F1	F2	F3	F4	F5	F6
TH (racemic)	-172.7	-118.8	-118.8	/	/	-212.2
TH (meso)	-175.0	-116.2	-122.2	/	/	-211.8
TH (meso)	-175.2	-116.6	-122.9	/	/	-212.4
HH (Figure S25)	175.1	-116.7	-123.4	-121.1	-127.1	-209.8
HH (Figure S26)	175.1	-116.7	-121.9	-117.3	-126.5	-199.1
HH (Figure S27)	175.1	-166.8	-122.0	-120.5	-127	-210.1
HH (Figure S28)	175.1	-116.5	-123.1	-118.7	-126.9	-199.2
CF ₂ -XA						
	F1	F2	F3	F4	F5	F6
HT	-84.7	-89.6	-204.4	/	/	/
HT	-84.5	-89.7	-204.2	/	/	/
HT	-84.9	-89.6	-205.2	/	/	/
HT	-85.0	-91.4	-201.2	/	/	/
TT	-84.5	-90.6	-206.5	-214.7	/	/
TT	-84.5	-90.6	-206.5	-214.3	/	/
TT	-84.6	-90.3	-206.1	-213.3	/	/
TT	-84.6	-90.3	-206.0	-213.3	/	/
TT	-84.7	-89.6	-212.7	-219.0	/	/
TT	-84.9	-98.8	-212.5	-219.5	/	/

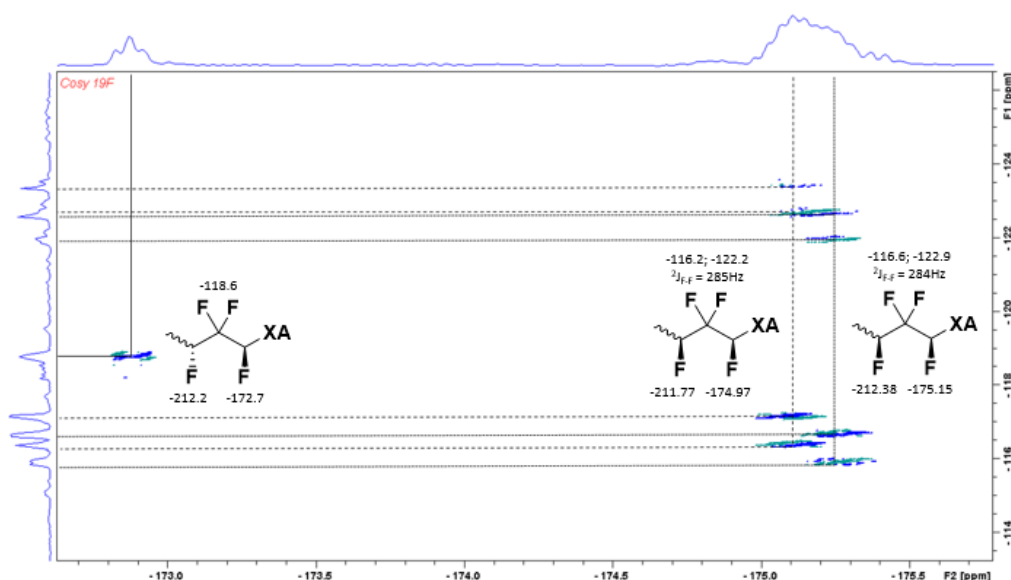


Figure 6. Zoom of the [-172.6 — -175.8ppm; -113.5 — -126.5ppm] region of the $^{19}\text{F}\{^1\text{H}\}$ COSY spectrum of PTrFE made by RAFT (Entry 4, Table 1) showing the CFH/CF₂ correlations of the [TH]-XA chain end.

allowed the identification of the corresponding fluorine and carbon resonances as a multiplet between -130.2 and -134.6 ppm and a singlet at 110.7 ppm in the ^{19}F and ^{13}C NMR spectra, respectively. The PTrFE-CFH₂ chain end ^{19}F NMR resonance, however, has not been previously assigned. The weak resonance observed at -244.6 ppm is assigned to this chain end in agreement with Dolbier,^{26,27} who reported a similar chemical shift for the -CFH₂ group of 1,1,2-trifluoroethane. The corresponding proton was found at 4.98 ppm (doublet of triplets, $^2J_{\text{H-F}} = 45.9$ Hz, $^3J_{\text{H-F}} = 12.7$ Hz).

The dormant chains, i.e. the xanthate terminated chain, produce very complex signals because of the chain inversions occurring during the polymerisation of TrFE.

To identify these chain ends, the ^{19}F NMR spectra of PTrFE synthesized by conventional radical and RAFT polymerisations (Figure S24) are compared. The signals ranging between -172 and -176 ppm are assigned to the -CFH-XA group while those between -84.2 and -91.5 ppm to the -CF₂-XA group.

In each case, several relatively broad peaks can be observed. Other sharp resonances are due to the monoadducts that have been assigned above. The multiplicity of the signals is likely caused by the chain defects occurring during the polymerisation.

C.3.1 Determination of -CFH-XA chain end

The ultimate CFH group (linked to the xanthate moiety) displays a number of correlations with both a CF₂ group and another CFH groups (around -120 and -210 ppm, respectively). Depending on how the ultimate TrFE unit was added to the polymer (via HH or TH additions), two different correlation systems were observed. The different -CFH-XA-terminated chains are detailed in Figure S25. For a better understanding, schematic representations of the two correlation systems are presented in Figure S26 (for the [TH]-XA, two splitting patterns are presented) and Figure S27 (for the [HH]-XA, only one splitting pattern is supplied).

First, the most intense correlations in the ^{19}F COSY spectrum are assigned to the [TH]-XA (-CF₂-CFH-CF₂-CFH-XA) chain end. Two different correlation systems were identified (Figure 6). In one system, the fluorine atom in the CFH resonance centered at -172.7 ppm couples with two isochronous fluorine atoms of a CF₂ group yielding the resonance centered at -118.6 ppm (no $^2J_{\text{F-F}}$ coupling nor difference of chemical shift are observed for both these fluorine atoms).

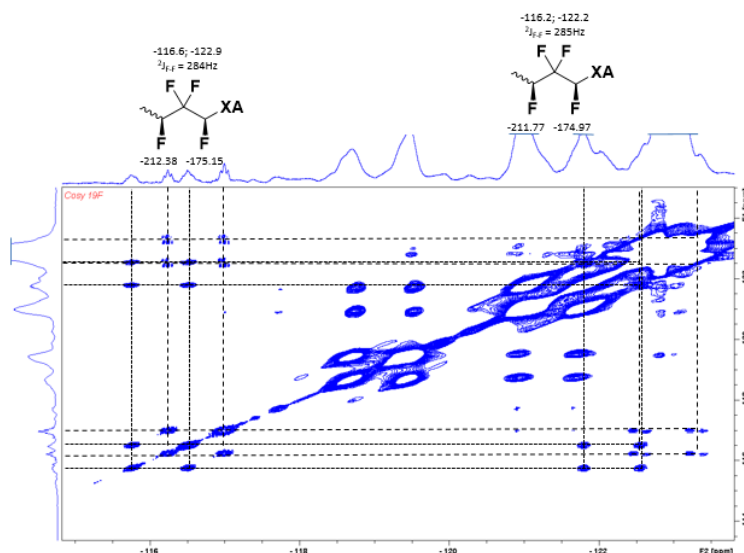


Figure 7. Zoom [-115.6 — -122.8ppm; -113.8 — 123.4ppm] on the ^{19}F COSY spectrum on the CF₂/CF₂ correlations zone for PTrFE made by RAFT (Entry 4, Table 1).

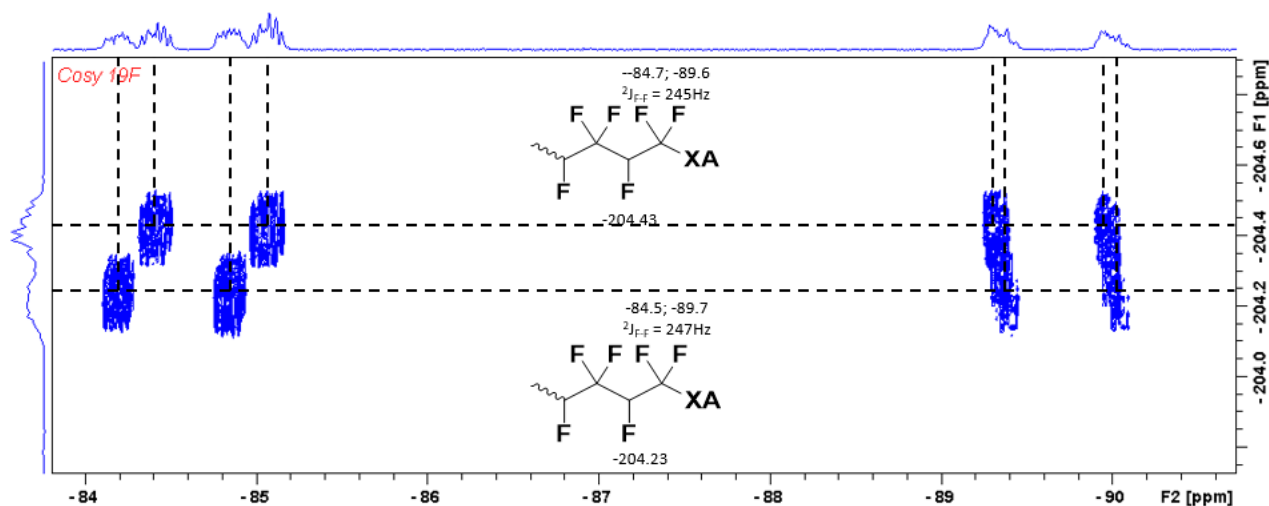


Figure 8. Zoom on the [-84 — -90.6ppm; -203.8 — 204.8ppm] region (CF_2/CFH correlations zone) of the ^{19}F COSY spectrum of a PTrFE made by RAFT (Entry 4, Table 1).

In addition, the fluorine atoms of this CF_2 group are coupled with another fluorine atom of a CFH group resonating at -212.2 ppm, confirming the expected [TH]-XA structure ($-\text{CF}_2\text{-CFH-CF}_2\text{-CFH-XA}$). The second [TH]-XA system is more complex (see schematic representation in Figure S26) with two sets of signals caused by the stereochemical configuration of the penultimate dyad. The fluorine atom (resonances at -175.15 and -175.97 ppm) of the terminal CFH group is coupled ($^3J_{\text{FF}}$ ranging from 10 to 20 Hz) with two non-equivalent fluorine atoms of the adjacent CF_2 group with resonances at -116.6, -122.9, and -116.2 and -122.2 ppm. This constitutes a meso dyad ($^2J_{\text{FF}} = 285$ Hz). A gHSQC experiment confirmed that these non-equivalent fluorine atoms are connected to the same carbon atom (resonances at 114.35 and 114.48 ppm respectively, Figure S28). The ^{19}F COSY spectrum does not show any other correlation than the 2J coupling in the $\text{CF}_2\text{-CF}_2$ correlation region (-115 to -123 ppm) (Figure 7). However, a 3-bond correlation (at -211.77 and -212.20 ppm) with the CFH of the penultimate TrFE unit was observed (Figure S29), confirming again the [TH]-XA expected structure. Interestingly, the 4J coupling between the ultimate and penultimate CFH group (Figure S30) could be observed on the COSY ^{19}F spectrum.

The less intense [HH]-XA correlation system ($-\text{CFH-CF}_2\text{-CF}_2\text{-CFH-XA}$) is assigned here. Depending on the configuration of the CFH groups, 4 different sets of signals have been observed and are summarised in Table 3 and in Figures S31 to S34. This type of chain end shows a typical correlation pattern in the ^{19}F COSY spectrum, between the fluorine atom of a CFH group (resonance around -175 ppm) and two sets of two non-equivalent fluorine atoms belonging to the CF_2 groups of the ultimate and penultimate TrFE units (Figures S31 to S34) resonating at around -116, -120, -123 and -127 ppm. This typical system is represented schematically in Figure S27. Each [HH]-XA chain end possesses the same splitting pattern, but with different chemical shifts due to the difference of stereochemical configuration of the CFH of the last three dyads.

As in the case of the [TH]-XA chain ends, each fluorine atom of this CF_2 groups yields a doublet of doublets with a large $^2J_{\text{F-F}}$ (ca. 282 Hz for the fluorine atoms at -116 and -123 ppm and around 287 Hz for the fluorine atoms at -120 and -127 ppm) and a smaller $^3J_{\text{F-F}}$ coupling constant of about 15-20 Hz (although an accurate measurement of this coupling constant was not possible). The values of these $^2J_{\text{FF}}$ coupling constants allow the identification of the fluorine atom resonance pairs at (-116, -123) and (-120, -126) ppm. This could not

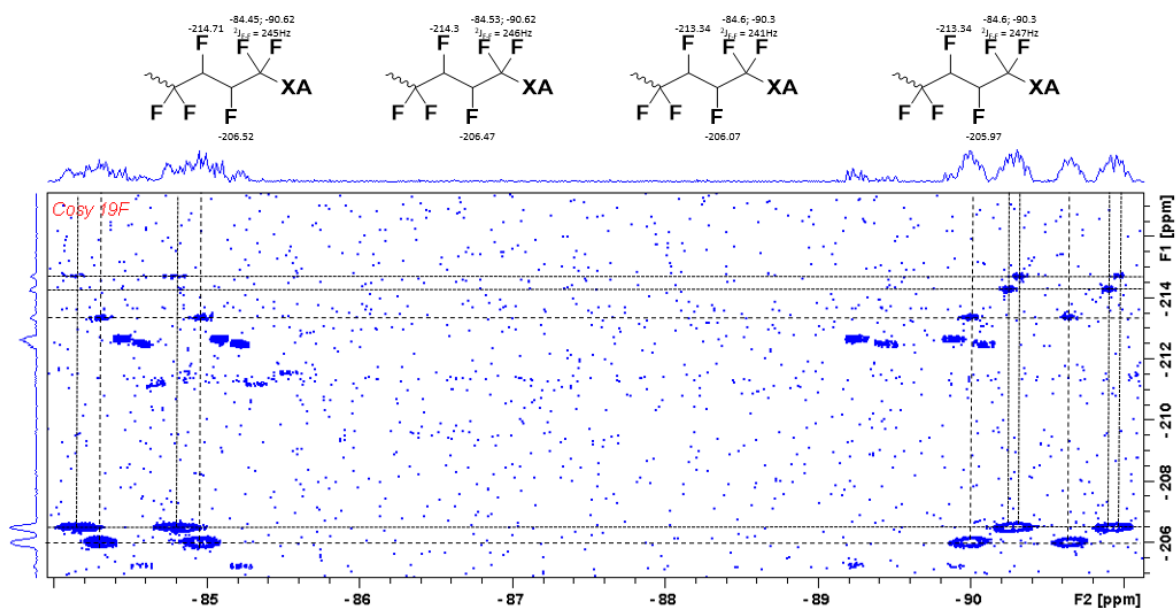


Figure 9. Zoom on the [-84 — -91ppm; -205 — -217ppm] region (CF_2/CFH correlations zone) of the ^{19}F COSY spectrum for PTrFE made by RAFT (Entry 4, Table 1).

be done using the gHSQC spectrum because the corresponding

Table 5. NMR signal assignments of the chain-ends of PtrFE made by RAFT polymerization.

	¹³ C	¹ H	¹⁹ F
α-chain-end and backbone			
-(CFH-CF ₂)-	116.1-117.7	n.a.	-110 - -140
-(CFH-CF ₂)-	83.4-87.1	5.4-6.1	-211.3 - -217.1
R-CH(CH ₃)-CFH-	12.0	1.4	n.a.
R-CH(CH ₃)-CFH-	38.7	3.2	n.a.
R-CH(CH ₃)-CFH- (RS/SR)	89.0;90.6; 90.6;90.5	5.2 (Broad peaks)	-196.1; -196.5; -197.0; -197.3
R-CH(CH ₃)-CFH- (RR/SS)	87.8;89.3; 89.5;89.2	5.3 (Broad peaks)	-208.5; -209.13; -209.3; -209.4
ω-chain end			
-CFH-XA	97.2 — 97.9	7.1-7.4	-171.1 — -175.6
-CFH-CF ₂ -CFH-XA (TH-XA, racemic)	n.d.	n.d.	-172.7
-CFH-CF ₂ -CFH-XA (TH-XA, racemic)	n.d.	n.a.	-118.6
-CFH-CF ₂ -CFH-XA (TH-XA, racemic)	n.d.	n.d.	-212.2
-CFH-CF ₂ -CFH-XA (TH-XA, meso 1)	n.d.	n.a.	-174.97
-CFH-CF ₂ -CFH-XA (TH-XA, meso 1)	114.48	n.a.	-116.2; -122.2
-CFH-CF ₂ -CFH-XA (TH-XA, meso 1)	n.d.	n.d.	-211.8
-CFH-CF ₂ -CFH-XA (TH-XA, meso 2)	n.d.	n.d.	-175.2
-CFH-CF ₂ -CFH-XA (TH-XA, meso 2)	114.35	n.a.	-116.6; -122.9
-CFH-CF ₂ -CFH-XA (TH-XA, meso 2)	n.d.	n.d.	-212.4
-CFH-CF ₂ -CFH-XA (TH-XA, meso)			
CFH-CF ₂ -CF ₂ -CFH-XA ([HH]-XA)			
CF ₂ -XA	123.5	n.a.	-84.1 to -91.3
Mono adducts			
R-CFH-CF ₂ -XA (SS/RR)	n.d.	5.6	-203.4 (dd, ³ J _{F-F} = 16 Hz; 16.4 Hz)
R-CFH-CF ₂ -XA (RS/SR)	n.d.	5.3	-189.2 (dd, ³ J _{F-F} = 17.3 Hz; 15.1 Hz)
R-CFH-CF ₂ -XA (SS/RR)	n.d.	n.a.	-87.3; -89.9 (dd, ² J _{F-F} = 237 Hz; ³ J _{F-F} = 16 Hz; 16.4 Hz)
R-CFH-CF ₂ -XA (RS/SR)	n.d.	n.a.	-82.1; -89.8 (dd, ² J _{F-F} = 236 Hz; ³ J _{F-F} = 17.2 Hz; 15.5 Hz)
R-CF ₂ -CFH-XA (SR/RS)	n.d.	n.a.	-108.8; -112.3 (dd, ³ J _{F-F} = 17.2 Hz/15.5 Hz; ² J _{F-F} = 260 Hz)
R-CF ₂ -CFH-XA (RR/SS)	n.d.	n.a.	-109.3; -113.1 (dd, ³ J _{F-F} = 17.7 Hz/16.9 Hz; ² J _{F-F} = 263.5 Hz)
R-CF ₂ -CFH-XA (SR/RS)	n.d.	7.2	-171.1 (dd, ³ J _{F-F} = 17.2 Hz/15.5 Hz)
R-CF ₂ -CFH-XA (RR/SS)	n.d.	7.1	-173.8 (dd, ³ J _{F-F} = 17.7 Hz/16.9 Hz)
Dead chain end			
-CF ₂ H	110.7	6.5 (broad triplet)	-130.4 to -134.4
-CFH ₂	77.8	5.0	-244.6

signals were too weak. All these CF₂ fluorine atoms show correlation spots with each other, further confirming the -CF₂-CF₂-CFH-XA sequence (Figure S35). The coupling with the fluorine atom of the CFH group (resonance at -199 ppm or -210 ppm depending of its stereochemical configuration) of the penultimate TrFE unit is also visible in the ¹⁹F COSY spectrum (Figure S36).

C.3.2 Determination of -CF₂-XA chain end

The CF₂-XA zone (from -84 to -92 ppm) displays 10 different structures corresponding to the [HT]-XA and [TT]-XA chain ends. The most representative correlations on the ¹⁹F COSY spectrum are presented Figure 8 and Figure 9. The other correlations are barely visible because they are too close to the background signal. The broad resonances visible in the ¹⁹F{¹H} spectrum centered at -84.2

and -91.5 ppm stem from the overlap of the ultimate CF₂ group signals. The non-equivalence of the fluorine atoms of this ultimate CF₂ moiety results in the presence of two sets of split resonances, with a strong ²J_{F-F} coupling of about 245 Hz. As in the previous cases of the CFH-XA chain ends, the lack of resolution does not allow an accurate determination of ³J_{F-F}. Among the 10 correlation systems observed, four CF₂ groups are coupled with only one CFH and are thus assigned to the regular HT structure (-CFH-CF₂-CFH-CF₂-XA, Figure 8). The structures where the ultimate CF₂ group is coupled with two CFH groups were attributed to the [TT]-XA structure (-CF₂-CFH-CFH-CF₂-XA, Figure 9). The ¹⁹F COSY spectrum does not allow the accurate assignment of the CF₂ group of the penultimate TrFE unit due to the important number of weak signals in the same area of the spectrum. This complexity arises from the large number of structures and complex splitting patterns resulting from the CFH stereocenters and the various chain defects possibly present close to the penultimate TrFE unit. Indeed, the chemical shifts will be slightly different for a [HT]-XA (or [TT]-XA) chain end if an HHTT addition (or HH addition) occurred a couple of TrFE units before, making the complete and accurate assignment of the -CF₂-XA chain end impossible.

Table 4 summarises the chemical shifts and assignments of the ¹H and ¹⁹F NMR signals of these PTrFE ω-chain end. Table 5 summarises the chemical shifts and assignments of the ¹H and ¹⁹F NMR signals of both the α- and ω-chain end of PTrFE made by RAFT polymerisation.

Conclusion

Using a large array of NMR experiments combining the three polymer nuclei (¹H, ¹³C and ¹⁹F), the chain-end microstructures of a PTrFE made by RAFT have been fully described. The study of the α-chain-end and of the monoadducts formed at the very beginning of the polymerisation showed that, unexpectedly, the initiation occurs preferentially on the head (CF₂ group) of the monomer rather than its tail (CFH group). This oddity has been rationalized by DFT calculations as resulting from an H-bond stabilization of the transition state and product of the head-end addition. The study of the ω-chain end allowed the unequivocal identification of the various termini formed by irreversible transfers (-CF₂H, and CFH₂) or by reversible transfer to the RAFT chain transfer agent (CF₂-XA, and CFH-XA, where XA is a O-ethyl xanthate group). This work allows a better understanding of the RAFT polymerisation of TrFE and paves the way to the study of the RAFT copolymerisation of VDF and TrFE.

Author Contributions

Vincent BOUAD: Conceptualization, Methodology, validation, Investigation, Writing Original Draft, Review & Editing. Marc GUERRE: Conceptualization, Methodology. Cedric TOTÉE: Methodology. Gilles SILLY: Methodology. Bruno AMEDURI: Writing Review & Editing. Rinaldo POLI: Methodology, Software, Formal analysis, Writing - Review & Editing. Vincent LADMIRAL: Conceptualization, Methodology, Writing Original draft, Review & Editing, Supervision, Funding acquisition.

Conflicts of interest

There are no conflicts to declare.

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